

**ตาราง C : ค่าศักย์ไฟฟ้ารีดักชันมาตรฐาน (ที่ 25°C)**

ครึ่งปฏิกิริยา	$E^0$ (volts)	$E^0$ (volts)	เงื่อนไขสำหรับ $E^0$
$F_2+2H^++2e \rightleftharpoons 2HF$	+3.06		
$F_2+2e \rightleftharpoons 2F^-$	+2.85		
$S_2O_8^{2-}+2e \rightleftharpoons 2SO_4^{2-}$	+2.01		
$Co^{3+}+e \rightleftharpoons Co^{2+}$	+1.82		
$H_2O_2+2H^++2e \rightleftharpoons 2H_2O$	+1.77		
$MnO_4^-+4H^++3e \rightleftharpoons MnO_2+2H_2O$	+ 1.695		
$Ce^{4+}+e \rightleftharpoons Ce^{3+}$		+1.70	1M HClO <sub>4</sub>
		+1.61	1M HNO <sub>3</sub>
		+1.44	1M H <sub>2</sub> SO <sub>4</sub>
		+1.28	1M HCl
$2HClO+2H^++2e \rightleftharpoons Cl_2+2H_2O$	+1.63		
$NaBiO_3+6H^++2e \rightleftharpoons Na^++Bi^{3+}+3H_2O$	+1.6		
$H_5IO_6+H^++2e \rightleftharpoons IO_3^-+3H_2O$	+1.6		
$2BrO_3^-+12H^++10e \rightleftharpoons Br_2+6H_2O$	+1.52		
$MnO_4^-+8H^++5e \rightleftharpoons Mn^{2+}+4H_2O$	+1.51		
$Mn^{3+}+e \rightleftharpoons Mn^{2+}$	+1.51		
$PbO_2+4H^++2e \rightleftharpoons Pb^{2+}+2H_2O$	1.455		
$Cl_2+2e \rightleftharpoons 2Cl^-$	+1.359		
$Cr_2O_7^{2-}+14H^++6e \rightleftharpoons 2Cr^{3+}+7H_2O$	+1.33		
		+1.09	1M HCl
$MnO_2+4H^++2e \rightleftharpoons Mn^{2+}+2H_2O$	+1.23		
		+1.24	1M HClO <sub>4</sub>
$O_2+4H^++4e \rightleftharpoons 2H_2O$	+1.229		
$2IO_3^-+12H^++10e \rightleftharpoons I_2+6H_2O$	+1.195		

ครึ่งปฏิกิริยา	$E^0$ (volts)	$E^0$ (volts)	เงื่อนไขสำหรับ $E^0$
$\text{Br}_{2(\text{aq})} + 2e \rightleftharpoons 2\text{Br}^-$	+1.087 <sup>a</sup>		
$\text{Br}_2(l) + 2e \rightleftharpoons 2\text{Br}^-$	+1.065 <sup>a</sup>		
$2\text{ICl}_2^- + 2e \rightleftharpoons \text{I}_2 + 4\text{Cl}^-$	+1.06		
$\text{HNO}_2 + \text{H}^+ + e \rightleftharpoons \text{NO} + \text{H}_2\text{O}$	+1.00		
$\text{NO}_3^- + 4\text{H}^+ + 3e \rightleftharpoons \text{NO} + 2\text{H}_2\text{O}$	+0.96		
$\text{NO}_3^- + 3\text{H}^+ + 2e \rightleftharpoons \text{HNO}_2 + \text{H}_2\text{O}$	+0.94		
$\text{NO}_3^- + 10\text{H}^+ + 8e \rightleftharpoons \text{NH}_4^+ + 3\text{H}_2\text{O}$	+0.87		
$2\text{Hg}^{2+} + 2e \rightleftharpoons \text{Hg}_2^{2+}$	+0.920		
$\text{Cu}^{2+} + \text{I}^- + e \rightleftharpoons \text{CuI}$	+0.86		
$\text{Ag}^+ + e \rightleftharpoons \text{Ag}$	+0.7994		
$\text{Hg}_2^{2+} + 2e \rightleftharpoons 2\text{Hg}$	+0.789		
$\text{Fe}^{3+} + e \rightleftharpoons \text{Fe}^{2+}$	+0.771		
		+0.732	1M $\text{HClO}_4$
		+0.700	1M $\text{HCl}$
		+0.674	1M $\text{H}_2\text{SO}_4$
		+0.46	2M $\text{H}_3\text{PO}_4$
$\text{O}_2 + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{O}_2$	+0.682		
$\text{I}_2(\text{aq}) + 2e \rightleftharpoons 2\text{I}^-$	+0.6197 <sup>b</sup>		
$\text{MnO}_4^- + e \rightleftharpoons \text{MnO}_4^{2-}$	+0.564		
$\text{H}_3\text{AsO}_4 + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_3\text{AsO}_3 + \text{H}_2\text{O}$	+0.559		
		+0.577	1M $\text{HCl}$ , 1M $\text{HClO}_4$
$(\text{I}_3^-) + 2e \rightleftharpoons 3\text{I}^-$	+0.536		
$\text{I}_2 + 2e \rightleftharpoons 2\text{I}^-$	+0.5355 <sup>b</sup>		
$\text{Cu}^+ + e \rightleftharpoons \text{Cu}$	+0.521		
$\text{H}_2\text{SO}_3 + 4\text{H}^+ + 4e \rightleftharpoons \text{S} + 3\text{H}_2\text{O}$	+0.45		
$\text{Ag}_2\text{CrO}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{CrO}_4^{2-}$	+0.446		
$2\text{H}_2\text{SO}_3 + 2\text{H}^+ + 4e \rightleftharpoons \text{S}_2\text{O}_3^{2-} + 3\text{H}_2\text{O}$	+0.40		

ครึ่งปฏิกิริยา	$E^0$ (volts)	$E^0$ (volts)	เงื่อนไขสำหรับ $E^0$
$[\text{Fe}(\text{CN})_6]^{3-} + e \rightleftharpoons [\text{Fe}(\text{CN})_6]^{4-}$	+0.356		
		+0.71	1M HCl
$\text{Cu}^{2+} + 2e \rightleftharpoons \text{Cu}$	+0.337		
$\text{Hg}_2\text{Cl}_2 + 2e \rightleftharpoons 2\text{Hg} + 2\text{Cl}^-$	+0.2680		
		+0.3337	0.1M KCl
		+0.2801	1M KCl
		+0.2412	KCl satur.
$\text{AgCl} + e \rightleftharpoons \text{Ag} + \text{Cl}^-$	+0.2224		
$\text{SO}_4^{2-} + 4\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{SO}_4 + 2\text{H}_2\text{O}$	+0.17		
$\text{Cu}^{2+} + e \rightleftharpoons \text{Cu}^+$	+0.153		
$\text{S} + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{S}$	+0.141		
$\text{Sn}^{4+} + 2e \rightleftharpoons \text{Sn}^{2+}$	+0.15		
$\text{AgBr} + e \rightleftharpoons \text{Ag} + \text{Br}^-$	+0.095		
$[\text{SnCl}_6]^{2-} + 2e \rightleftharpoons [\text{SnCl}_4]^{2-} + 2\text{Cl}^-$		+0.14	1M HCl
$\text{S}_4\text{O}_6^{2-} + 2e \rightleftharpoons 2\text{S}_2\text{O}_3^{2-}$	0.08		
$[\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-} + e \rightleftharpoons \text{Ag} + 2\text{S}_2\text{O}_3^{2-}$	+0.01		
$2\text{H}^+ + 2e \rightleftharpoons \text{H}_2$	0.000		
$\text{Pb}^{2+} + 2e \rightleftharpoons \text{Pb}$	-0.126		
$\text{Sn}^{2+} + 2e \rightleftharpoons \text{Sn}$	-0.136		
$\text{AgI} + e \rightleftharpoons \text{Ag} + \text{I}^-$	-0.151		
$\text{CuI} + e \rightleftharpoons \text{Cu} + \text{I}^-$	-0.185		
$\text{Ni}^{2+} + 2e \rightleftharpoons \text{Ni}$	-0.250		
$\text{V}^3 + e \rightleftharpoons \text{V}^{2+}$	-0.255		
$\text{CO}^{2+} + 2e \rightleftharpoons \text{Co}$	-0.277		
$[\text{Ag}(\text{CN})_2]^- + e \rightleftharpoons \text{Ag} + 2\text{CN}^-$	-0.31		
$\text{PbSO}_4 + 2e \rightleftharpoons \text{Pb} + \text{SO}_4^{2-}$	-0.356		

ครึ่งปฏิกิริยา	$E^0$ (volts)	$E^{0'}$ (volts)	เงื่อนไขสำหรับ $E^{0'}$
$\text{Cd}^{2+} + 2e \rightleftharpoons \text{Cd}$	-0.403		
$\text{Cr}^{3+} + e \rightleftharpoons \text{Cr}^{2+}$	-0.41		
$\text{Fe}^{2+} + 2e \rightleftharpoons \text{Fe}$	-0.440		
$\text{Cr}^{3+} + 3e \rightleftharpoons \text{Cr}$	-0.74		
$\text{Zn}^{2+} + 2e \rightleftharpoons \text{Zn}$	-0.763		
$\text{Mn}^{2+} + 2e \rightleftharpoons \text{Mn}$	-1.18		
$\text{Al}^{3+} + 3e \rightleftharpoons \text{Al}$	-1.66		
$\text{Mg}^{2+} + 2e \rightleftharpoons \text{Mg}$	-2.37		
$\text{Na}^+ + e \rightleftharpoons \text{Na}$	-2.71		
$\text{Ca}^{2+} + 2e \rightleftharpoons \text{Ca}$	-2.87		
$\text{Sr}^{2+} + 2e \rightleftharpoons \text{Sr}$	-2.89		
$\text{Ba}^{2+} + 2e \rightleftharpoons \text{Ba}$	-2.90		
$\text{K}^+ + e \rightleftharpoons \text{K}$	-2.92		
$\text{Li}^+ + e \rightleftharpoons \text{Li}$	-3.04		
2. Alkali Solutions			
$\text{ClO}^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{Cl}^- + 2\text{OH}^-$	+0.89		
$\text{O}_2^{2-} + 2\text{H}_2\text{O} + 2e \rightleftharpoons 4\text{OH}^-$	+0.88		
$\text{BrO}^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{Br}^- + 2\text{OH}^-$	+0.76		
$\text{MnO}_4^- + 2\text{H}_2\text{O} + 3e \rightleftharpoons \text{MnO}_2 + 4\text{OH}^-$	+0.59		
$\text{O}_2 + 2\text{H}_2\text{O} + 4e \rightleftharpoons 4\text{OH}^-$	+0.401		
$[\text{Ag}(\text{NH}_3)_2]^+ + e \rightleftharpoons \text{Ag} + 2\text{NH}_3$	+0.373		
$\text{ClO}_3^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{ClO}_2^- + 2\text{OH}^-$	+0.33		
$\text{Co}(\text{OH})_3 + e \rightleftharpoons \text{Co}(\text{OH})_2 + \text{OH}^-$	+0.17		
$[\text{Co}(\text{NH}_3)_6]^{3+} + e \rightleftharpoons [\text{Co}(\text{NH}_3)_6]^{2+}$	+0.1		
$\text{NO}_3^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{NO}_2^- + 2\text{OH}^-$	+0.01		
$\text{MnO}_2 + 2\text{H}_2\text{O} + 2e \rightleftharpoons \text{Mn}(\text{OH})_2 + 2\text{OH}^-$	-0.05		
$[\text{Cu}(\text{NH}_3)_4]^{2+} + 2e \rightleftharpoons \text{Cu} + 4\text{NH}_3$	-0.11		

ครึ่งปฏิกิริยา	$E^0$ (volts)	$E^{0'}$ (volts)	เงื่อนไขสำหรับ $E^0$
$\text{NO}_3^- + 6\text{H}_2\text{O} + 8e \rightleftharpoons \text{NH}_3 + 9\text{OH}^-$	-0.13		
$\text{NO}_2^- + 5\text{H}_2\text{O} + 6e \rightleftharpoons \text{NH}_3 + 7\text{OH}^-$	-0.18		
$[\text{Ag}(\text{CN})_2]^- + e \rightleftharpoons \text{Ag} + 2\text{CN}^-$	-0.31		
$\text{NO}_2 + \text{H}_2\text{O} + e \rightleftharpoons 2\text{OH}^- + \text{NO}$	-0.46		
$\text{S} + 2e \rightleftharpoons \text{S}^{2-}$	-0.48		
$[\text{Pb}(\text{OH})_3]^- + 2e \rightleftharpoons 3\text{OH}^-$	-0.54		
$\text{Fe}(\text{OH})_3 + e \rightleftharpoons \text{Fe}(\text{OH})_2 + \text{OH}^-$	-0.56		
$[\text{Cd}(\text{NH}_3)_4]^{2+} + 2e \rightleftharpoons \text{Cd} + 4\text{NH}_3$	-0.597		
$\text{AsO}_4^{3-} + \text{H}_2\text{O} + 2e \rightleftharpoons \text{AsO}_3^{3-} + 2\text{OH}^-$	-0.67		
$\text{HgS} + 2e \rightleftharpoons \text{Hg} + \text{S}^{2-}$	-0.72		
$2\text{H}_2\text{O} + 2e \rightleftharpoons \text{H}_2 + 2\text{OH}^-$	-0.828		
$[\text{Sn}(\text{OH})_6]^{2-} + 2e \rightleftharpoons [\text{Sn}(\text{OH})_3]^- + 3\text{OH}^-$	-0.90		
$[\text{Zn}(\text{NH}_3)_4]^{2+} + 2e \rightleftharpoons \text{Zn} + 4\text{NH}_3$	-1.03		
$[\text{Cd}(\text{CN})_4]^{2-} + 2e \rightleftharpoons \text{Cd} + 4\text{CN}^-$	-1.03		
$[\text{Cr}(\text{OH})_4]^- + 3e \rightleftharpoons \text{Cr} + 4\text{OH}^-$	-1.2		
$[\text{Zn}(\text{OH})_4]^{2-} + 2e \rightleftharpoons \text{Zn} + 4\text{OH}^-$	-1.22		
$\text{ZnS} + 2e \rightleftharpoons \text{Zn} + \text{S}^{2-}$	-1.44		
$\text{Mn}(\text{OH})_2 + 2e \rightleftharpoons \text{Mn} + 2\text{OH}^-$	-1.55		
$[\text{Al}(\text{OH})_4]^- + 3e \rightleftharpoons \text{Al} + 4\text{OH}^-$	-2.35		