ภาคผนวก B

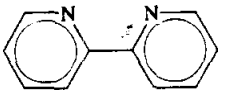
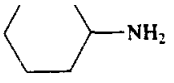
ตาราง1B : ค่าคงที่ของการแตกตัวของกรด (Acid Dissocistion Constants)

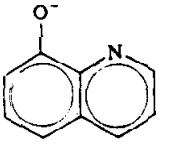
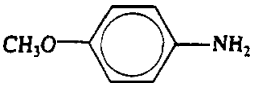
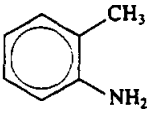
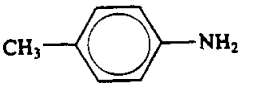
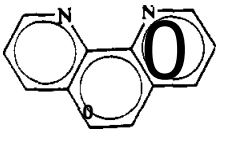
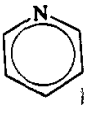
ชื่อ	สูตร	K_a	pK_a
Acetic Acid	CH_3CO_2H	1.76×10^{-5}	4.754
Adipic Acid	$HO_2C(CH_2)_4CO_2H$	3.8×10^{-5}	4.42
		3.8×10^{-16}	5.42
Arsenic Acid	H_3AsO_4	5.8×10^{-3}	2.24
		1.1×10^{-7}	6.96
		3.2×10^{-12}	1 1.49
Arsenious Acid	H_3AsO_3	5.1×10^{-10}	9.29
Benzoic Acid	$C_6H_5CO_2H$	6.28×10^{-5}	4.202
Boric Acid	H_3BO_3	5.81×10^{-10}	9.236
		1.8×10^{-13}	12.74
		1.6×10^{-14}	13.80
Carbonic Acid	H_2CO_3	4.45×10^{-7}	6.352
		4.69×10^{-11}	10.329
Chloroacetic Acid	$ClCH_2CO_2H$	1.36×10^{-3}	2.866
Chlorous Acid	$HClO_2$	1.1×10^{-2}	1.96
Citric Acid	CO_2H	7.45×10^{-4}	3.128
	$HO_2CCH_2CCH_2CO_2H$	1.73×10^{-5}	4.762
	OH	4.02×10^{-7}	6.396
Dichloroacetic Acid	Cl_2CHCO_2H	5.0×10^{-2}	1.30
Ethylenediaminetetra	$(HO_2CCH_2)_2NCH_2CH_2N$	1×10^{-2}	2.0
acetic Acid	$(CH_2CO_2H)_2$	2.1×10^{-3}	2.68
		7.8×10^{-7}	6.1 1
		6.8×10^{-11}	10.17
Formic Acid	HCO_2H	1.80×10^{-4}	3.745
Fumaric Acid	HO_2C H	8.85×10^{-4}	3.053
	c=c'	3.21×10^{-5}	4.493
	H' CO_2H		

ชื่อ	สูตร	K_a	pK_a
Hydrazoic Acid	HN_3	2.2×10^{-5}	4.66
Hydrocyanic Acid	HCN	6.2×10^{-10}	9.21
Hydrofluoric Acid	HF	6.8×10^{-4}	3.17
Hydrogen peroxide	H_2O_2	2.2×10^{-12}	11.66
Hydrosulphuric Acid	H_2S	9.5×10^{-8}	7.02
		1×10^{-14}	14.0
Hydrothiocyanic Acid	HSCN	1×10^{-1}	1.0
Hydroacetic acid Acid	$HOCH_2CO_2H$	1.48×10^{-4}	3.830
Hypochlorous Acid	HClO	3.0×10^{-8}	7.52
Hypophosphorous Acid	$HOPH_2$ (O)	5.9×10^{-2}	1.23
Iodic Acid	$HOI{=}O$ (O)	1.7×10^{-1}	0.77
Lactic Acid	CH_3CHCO_2H (OH)	1.35×10^{-4}	3.870
Maleic Acid	$HO_2C(H)C{=}C(H)CO_2H$	1.23×10^{-2}	1.910
		4.66×10^{-7}	6.332
Malic Acid	$HO_2CCH_2CHCO_2H$ (OH)	3.48×10^{-4}	3.458
		8.00×10^{-6}	5.097
Malonic Acid	$HO_2CCH_2CO_2H$	1.42×10^{-3}	2.848
		2.01×10^{-6}	5.697
Nitrilotriacetic Acid	$N(CH_2CO_2H)_3$	2.24×10^{-2}	1.650
		1.15×10^{-3}	2.939
		4.63×10^{-11}	10.334
Nitrous Acid	HNO_2	7.1×10^{-4}	3.15
Oxalic Acid	HO_2CCO_2H	5.60×10^{-2}	1.252
		5.42×10^{-5}	4.266
Periodic Acid	H_5IO_6	2.6×10^{-2}	1.59
		5.1×10^{-9}	8.29

ชื่อ	สูตร	K_a	pK_a
Phenol	C_6H_5OH	1.05×10^{-10}	9.979
Phenylacetic Acid	$C_6H_5CH_2CO_2H$	4.90×10^{-5}	4.310
Phosphoric Acid	H_3PO_4	7.11×10^{-3}	2.148
		6.32×10^{-8}	7.199
		4.5×10^{-13}	12.35
Phosphorous Acid	$\overset{O}{\overset{\|\|}{HP(OH)_2}}$	3×10^{-2}	1.5
		1.6×10^{-7}	6.80
Phthalic Acid	$C_6H_4(CO_2H)_2$	1.12×10^{-3}	2.951
		3.91×10^{-6}	5.408
Salicylic Acid	$C_6H_4(CO_2H)(OH)$	1.1×10^{-3}	2.96
		1.0×10^{-14}	13.74
Succinic Acid	$HO_2CCH_2CH_2CO_2H$	6.21×10^{-5}	4.207
		2.31×10^{-6}	5.636
Sulfamic Acid	H_2NSO_3H	1.03×10^{-1}	0.987
Sulfuric Acid	H_2SO_4		
		1.0×10^{-2}	2.00
Sulfurous Acid	H_2SO_3	1.2×10^{-2}	1.92
		6.6×10^{-8}	7.18
Tartaric Acid	$HO_2C\overset{OH}{\overset{\|}{C}}H\overset{OH}{\overset{\|}{C}}HCO_2H$	9.20×10^{-4}	3.036
		4.31×10^{-5}	4.366
Trichloroacetic Acid	Cl_3CCO_2H	2.2×10^{-1}	0.66

ตาราง 2B : ค่าคงที่ของการแตกตัวของเบส (Base Dissocistion Constants)

ชื่อ	สูตร	K_b	pK_b
Ammonia	NH_3	1.75×10^{-5}	4.757
Aniline	$C_6H_5NH_2$	3.99×10^{-10}	9.399
Benzylamine	$C_6H_5CH_2NH_2$	2.2×10^{-5}	4.66
2,2'-Bipyridine		2.2×10^{-10}	9.66
Butylamine	$CH_3(CH_2)_3NH_2$	4.37×10^{-4}	3.360
Cyclohexylamine		4.4×10^{-4}	3.36
Diethylamine	$(CH_3CH_2)_2NH$	8.57×10^{-4}	3.067
Dimetylamine	$(CH_3)_2NH$	5.94×10^{-4}	3.226
Ethanolamine	$HOCH_2CH_2NH_2$	3.15×10^{-5}	4.502
Ethylamine	$CH_3CH_2NH_2$	4.33×10^{-4}	3.364
Ethylenediamine	$NH_2CH_2CH_2NH_2$	8.47×10^{-5}	4.072
		7.05×10^{-8}	7.152
Glycinate	$NH_2CH_2CO_2^-$	6.00×10^{-5}	4.222
		2.24×10^{-12}	11.650

ชื่อ	สูตร	K_b	pK_b
Hydrazine	H_2NNH_2	$9.5 \times 10^{-}$	6.02
Hydroxylamine	$HONH_2$	9.1×10^{-9}	8.04
8-Hydroxyquinoline		6.5×10^{-5}	4.19
		8.1×10^{-10}	9.09
4-Methoxyaniline (p-anisidine)		2.28×10^{-9}	8.642
Methylamine	CH_3NH_2	4.4×10^{-4}	3.36
2-Methoxyaniline (o-toluidine)		2.80×10^{-10}	9.553
4-Methoxyaniline (m-toluidine)		1.21×10^{-9}	8.917
1,10-Phenanthroline		7.2×10^{-10}	9.14
Propylamine	$CH_3CH_2CH_2NH_2$	3.68×10^{-4}	3.434
Pyridine		1.69×10^{-9}	8.772

ชื่อ	สูตร	K_b	pK_b
Triethanolamine	$(HOCH_2CH_2)_3N$	5.70×10^{-7}	6.238
Triethylamine	$(CH_3CH_2)_3N$	$5.19x\ 10^{-4}$	3.285
Trimethylamine	$(CH_3)_3N$	6.31×10^{-5}	4.200
Tris- (hydroxymethyl) aminomethane (Tris or THAM)	$(HOCH_2)_3CNH_2$	1.19×10^{-6}	5.924

ตาราง 3B : ค่าคงที่ของผลคูณการละลาย (Solubility Product constants, K_{sp})

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Acetates			
Silver acetate	$CH_3COOAg = Ag^+ + CH_3COO^-$	4×10^{-3}	2.4
Arsenates			
Silver arsenate	$Ag_3AsO_4 = 3Ag' + AsO_4^{3-}$	1×10^{-22}	22.0
Bromates			
Silver bromate	$AgBrO_3 = Ag^+ + BrO_3^-$	6×10^{-5}	4.2
Bromides			
Copper (I) bromide	$CuBr = Cu^{+2} + Br^-$	6×10^{-9}	8.2
Lead bromide	$PbBr_2 = Pb^{+2} + 2Br^-$	4.6×10^{-6}	5.34
Mercury (I) bromide	$Hg_2Br_2 = Hg_2^{+2} + 2Br^-$	1.3×10^{-22}	21.89
Silver bromide	$AgBr = Ag^+ + Br^-$	5×10^{-13}	12.3
Carbonates			
Barium carbonate	$BaCO_3 = Ba^{2+} + CO_3^{2-}$	6×10^{-9}	8.80
Cadmium carbonate	$CdCO_3 = Cd^{2+} + co,^{2-}$	5.2×10^{-12}	11.28
Calcium carbonate	$CaCO_3 = Ca^{2+} + CO_3^{2-}$	6.9×10^{-9}	8.16
Cobolt (II) carbonate	$CoCO_3 = Co^{2+} + CO_3^{2-}$	8×10^{-13}	12.1
Copper (II) carbonate	$CuCO_3 = Cu^{2+} + CO_3^{2-}$	2.5×10^{-10}	9.60
Iron (II) carbonate	$FeCO_3 = Fe^{2+} + CO_3^{2-}$	2.1×10^{-11}	10.68
Lead carbonate	$PbCO_3 = Pb^{2+} + CO_3^{2-}$	1.5×10^{-13}	12.82
Magnesium carbonate	$MgCO_3 = Mg^{2+} + CO_3^{2-}$	4×10^{-5}	4.4
Manganese (II)carbonate	$MnCO_3 = Mn^{2+} + CO_3^{2-}$	9×10^{-11}	10.1
Mercury (I) carbonate	$Hg_2CO_3 = 2Hg^+ + CO_3^{2-}$	9×10^{-17}	16.1
Nickel carbonate	$NiCO_3 = Ni^{2+} + CO_3^{2-}$	7×10^{-9}	8.2
Silver carbonate	$Ag_2CO_3 = 2Ag^+ + CO_3^{2-}$	8.2×10^{-12}	11.09
Strontium carbonate	$SrCO_3 = Sr^{2+} + CO_3^{2-}$	7×10^{-10}	9.2
Zinc carbonate	$ZnCO_3 = Zn^{2+} + CO_3^{2-}$	2×10^{-11}	10.7

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Chlorides			
Copper (I) chloride	$CuCl \rightleftharpoons Cu^{+} + Cl^{-}$	3.2×10^{-7}	6.49
Lead chloride	$PbCl_2 \rightleftharpoons Pb^{2+} + 2Cl^{-}$	1.6×10^{-5}	4.80
Mercury (I) chloride	$Hg_2Cl_2 \rightleftharpoons Hg_2^{2+} + 2Cl^{-}$	1.1×10^{-18}	17.96
Silver chloride	$AgCl \rightleftharpoons Ag^{+} + Cl^{-}$	1.8×10^{-10}	9.74
Thallium chloride	$TlCl \rightleftharpoons Tl^{+} + Cl^{-}$	3.5×10^{-4}	3.46
Chromates			
Barium chromates	$BaCrO_4 \rightleftharpoons Ba^{2+} + CrO_4^{2-}$	1.2×10^{-10}	9.92
Calcium chromates	$CaCrO_4 \rightleftharpoons Ca^{2+} + CrO_4^{2-}$	7.1×10^{-4}	3.15
Copper (I) chromates	$CuCrO_4 \rightleftharpoons Cu^{2+} + CrO_4^{2-}$	3.6×10^{-6}	5.44
Lead chromates	$PbCrO_4 \rightleftharpoons Pb^{2+} + CrO_4^{2-}$	2×10^{-14}	13.7
Mercury (I) chromates	$Hg_2CrO_4 \rightleftharpoons 2Hg^{+} + CrO_4^{2-}$	2×10^{-9}	8.7
Strontium chromates	$SrCrO_4 \rightleftharpoons Sr^{2+} + CrO_4^{2-}$	1.9×10^{-12}	11.72
Silver chromate	$Ag_2CrO_4 \rightleftharpoons 2Ag^{+} + CrO_4^{2-}$	3.6×10^{-5}	4.44
Cyanides			
Mercury (I) cyanide	$Hg_2(CN)_2 \rightleftharpoons Hg_2^{2+} + 2CN^{-}$	5×10^{-40}	39.3
Silver cyanide	$AgCN \rightleftharpoons Ag^{+} + CN^{-}$	1.6×10^{-14}	13.80
Ferrocyanlde			
Copper (II) ferrocyanides	$Cu_2[Fe(CN)_6] \rightleftharpoons 2Cu^{2+} + [Fe(CN)_6]^{4-}$	1.3×10^{-16}	15.89
Silver ferrocyanides	$Ag_4[Fe(CN)_6] \rightleftharpoons 4Ag^{2+} + [Fe(CN)_6]^{4-}$	1.6×10^{-41}	40.80
Zinc ferrocyanides	$K_2Zn_3[Fe(CN)_6]_2 \rightleftharpoons 2K^{+} + 3Zn^{2+} + 2[Fe(CN)_6]^{4-}$	1×10^{-95}	95.0
Fluorides			
Barium fluoride	$BaF_2 \rightleftharpoons Ba^{2+} + 2F^{-}$	2.4×10^{-5}	4.62
Calcium fluoride	$CaF_2 \rightleftharpoons Ca^{2+} + 2F^{-}$	1.7×10^{-10}	9.77
Lead fluoride	$PbF_2 \rightleftharpoons Pb^{2+} + 2F^{-}$	2.7×10^{-8}	7.57
Magnesium fluoride	$MgF_2 \rightleftharpoons Mg^{2+} + 2F^{-}$	6.5×10^{-9}	8.19

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Fluorides (ต่อ)			
Strontium fluoride	$SrF_2 = Sr^{2+} + 2F^-$	7.9×10^{-10}	9.10
Hydroxides			
Aluminium hydroxide	$Al(OH)_3 = Al^{3+} + 3OH^-$	5×10^{-33}	32.3
Cadmium hydroxide	$Cd(OH)_2 = Cd^{2+} + 2OH^-$	2.0×10^{-14}	13.70
Chromium (III)hydroxide	$Cr(OH)_3 = Cr^{3+} + 3OH^-$	7×10^{-31}	30.2
Cobalt (Ill) hydroxide	$Co(OH)_3 = Co^{3+} + 3OH^-$	2.5×10^{-43}	42.60
Cobalt (II) hydroxide	$Co(OH)_2 = Co^{2+} + 2OH^-$	2.5×10^{-16}	15.60
Copper (II) hydroxide	$Cu(OH)_2 = Cu^{2+} + 2OH^-$	1.6×10^{-19}	18.80
Iron (III) hydroxide	$Fe(OH)_3 = Fe^{3+} + 3OH^-$	6×10^{-38}	37.2
Iron (II) hydroxide	$Fe(OH)_2 = Fe^{2+} + 2OH^-$	1.8×10^{-15}	14.74
Lead hydroxide	$Pb(OH)_2 = Pb^{2+} + 2OH^-$	4.2×10^{-15}	14.38
Magnesium hydroxide	$Mg(OH)_2 = Mg^{2+} + 2OH^-$	8.9×10^{-12}	11.05
Manganese (I I) hydroxide	$Mn(OH)_2 = Mn^{2+} + 2OH^-$	2×10^{-13}	12.7
Mercury (II) hydroxide	$HgO + H_2O = Hg^{2+} + 2OH^-$	3×10^{-26}	25.5
Nickel hydroxide	$Ni(OH)_2 = Ni^{2+} + 2OH^-$	1.6×10^{-16}	15.80
Silver hydroxide	$1/2\ Ag_2O + 1/2\ H_2O = Ag^+ + OH^-$	2×10^{-8}	7.7
Tin (IV) hydroxide	$Sn(OH)_4 = Sn^{4+} + 4OH^-$	1×10^{-56}	56.0
Tin (II) hydroxide	$Sn(OH)_2 = Sn^{2+} + 2OH^-$	3×10^{-27}	26.5
Zinc hydroxide	$Zn(OH)_2 = Zn^{2+} + 2OH^-$	5×10^{-17}	16.3
Iodates			
Barium iodate	$Ba(IO_3)_2 = Ba^{2+} + 2\ IO_3^-$	1.3×10^{-9}	8.89
Calcium iodate	$Ca(IO_3)_2 = Ca^{2+} + 2\ IO_3^-$	1.7×10^{-6}	5.77
Lead iodate	$Pb(IO_3)_2 = Pb^{2+} + 2\ IO_3^-$	2.6×10^{-13}	12:59
Mercury (I) iodate	$Hg_2(IO_3)_2 = Hg_2^{2+} + 2\ IO_3^-$	1.9×10^{-14}	13.72
Mercury (II) iodate	$Hg(IO_3)_2 = Hg^{2+} + 2\ IO_3^-$	3×10^{-13}	12.5
Silver iodate	$AgIO_3 = Ag+ + IO_3^-$	3 x 10-s	7.5
Iodides			
Copper (I) iodide	$CuI = cu+ + I^-$	1×10^{-12}	12.0
Lead iodide	$PbI_2 = Pb^{2+} + 2I^-$	8.3×10^{-9}	8.08

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Iodides (ต่อ)			
Mercury (I) iodide	$Hg_2I_2 = Hg_2^{2+} + 2I^-$	4×10^{-29}	28.4
Mercury (II) iodide	$HgI_2 = Hg^{2+} + 2I^-$	4×10^{-18}	17.4
Silver iodide	$AgI = Ag^+ + I^-$	8.5×10^{-17}	16.07
Thallium(I) iodide	$TlI = Tl^+ + I^-$	2.5×10^{-8}	7.60
Nitrites			
Silver nitrite	$AgNO_2 = Ag^+ + NO_2^-$	1.2×10^{-4}	3.92
Oxalates			
Barium oxalate	$BaC_2O_4 = Ba^{2+} + C_2O_4^{2-}$	1.5×10^{-8}	7.82
Cadmium oxalate	$CdC_2O_4 = Cd^{2+} + C_2O_4^{2-}$	1.5×10^{-8}	7.82
Calcium oxalate	$CaC_2O_4 = Ca^{2+} + C_2O_4^{2-}$	1.3×10^{-9}	8.89
Iron (II) oxalate	$FeC_2O_4 = Fe^{2+} + C_2O_4^{2-}$	2×10^{-7}	6.7
Magnesium oxalate	$MgC_2O_4 = Mg^{2+} + C_2O_4^{2-}$	8.6×10^{-5}	4.07
Manganese (III) oxalate	$Mn_2(C_2O_4)_3 = 2Mn^{3+} + 3C_2O_4^{2-}$	7×10^{-20}	19.2
Manganese (II) oxalate	$MnC_2O_4 = Mn^{2+} + C_2O_4^{2-}$	1.1×10^{-15}	14.96
Silver oxalate	$Ag_2C_2O_4 = 2Ag^{2+} + C_2O_4^{2-}$	1×10^{-11}	11.0
Strontium oxalate	$SrC_2O_4 = Sr^{2+} + C_2O_4^{2-}$	5.6×10^{-8}	7.25
Zinc oxalate	$ZnC_2O_4 = Zn^{2+} + C_2O_4^{2-}$	1.5×10^{-9}	8.82
Phosphates			
Barium phosphate	$Ba_3(PO_4)_2 = 3Ba^{2+} + 2PO_4^{3-}$	6×10^{-39}	38.2
Calcium phosphate	$Ca_3(PO_4)_2 = 3Ca^{2+} + 2PO_4^{3-}$	1×10^{-25}	25.0
Iron (III) phosphate	$FePO_4 = Fe^{3+} + PO_4^{3-}$	1.3×10^{-22}	21.89
Magnesium Ammonium phosphate	$MgNH_4PO_4 = Mg^{2+} + + NH_4^+ + PO_4^{3-}$	2×10^{-13}	12.7
Magnesium phosphate	$Mg_3(PO_4)_2 = 3Mg^{2+} + 2PO_4^{3-}$	2.6×10^{-13}	12.59
Silver phosphate	$Ag_3PO_4 = 3Ag^+ + PO_4^{3-}$	1.8×10^{-18}	17.74
Strontium phosphate	$Sr_3(PO_4)_2 = 3Sr^{2+} + 2PO_4^{3-}$	1×10^{-31}	31.0
Zirconium phosphate	$Zr_3(PO_4)_4 = 3Zr^{4+} + 4PO_4^{3-}$	1×10^{-132}	132

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Sulfates			
Barium sulfate	$BaSO_4 = Ba^{2+} + SO_4^{2-}$	1.5×10^{-9}	8.82
Calcium sulfate	$CaSO_4 = Ca^{2+} + SO_4^{2-}$	2.4×10^{-5}	4.62
Lead sulfate	$PbSO_4 = Pb^{2+} + SO_4^{2-}$	1.3×10^{-8}	7.89
Mercury (I) sulfate	$Hg_2SO_4 = Hg_2^{2+} + SO_4^{2}$	1×10^{-6}	6.0
Silver sulfate	$Ag_2SO_4 = 2Ag^{+} + SO_4^{2-}$	1.6×10^{-5}	4.80
Strontium sulfate	$SrSO_4 = Sr^{2+} + SO_4^{2-}$	2.8×10^{-7}	6.55
Sulfides			
Bismuth sulfide	$Bi_2S_3 = 2Bi^{3+} + 3S^{2}$	1×10^{-100}	100
Cadmium sulfide	$CdS = Cd^{2+} + S^{2-}$	6×10^{-27}	26.2
Cobalt (II) sulfide	$cos = Co^{2+} + S^{2-}$	5×10^{-22} *(a)*	21.3
		6×10^{-29} (β)	28.2
Copper (I) sulfide	$Cu_2S = 2Cu^{+} + S^{2-}$	1.2×10^{-49}	48.92
Copper (II) sulfide	$cus = Cu^{2+} + S^{2-}$	4×10^{-36}	35.4
Iron (III) sulfide	$Fe_2S_3 = 2Fe^{3+} + 3S^{2-}$	1×10^{-88}	88.0
Iron (II) sulfide	$FeS = Fe^{2+} + S^{2-}$	5×10^{-18}	17.3
Lead sulfide	$PbS = Pb^{2+} + S^{2-}$	8×10^{-28}	27.1
Manganese (II) sulfide	$MnS = Mn^{2+} + S^{2-}$	8×10^{-14}	13.1
Mercury(I) sulfide	$Hg_2S = Hg_2^{2+} + S^{2-}$	1×10^{-45}	45.0
Mercyry (II) sulfide	$HgS = Hg^{2+} + S^{2-}$	1×10^{-50}	50.0
Nickel sulfide	$NiS = Ni^{2+} + S^{2-}$	1×10^{-22} (α)	22.0
		3×10^{-28} (β)	27.5
		7×10^{-30} (γ)	29.2
Silver sulfide	$Ag_2S = 2Ag^{+} + S^{2-}$	1×10^{-50}	50.0
Tin (II) sulfide	$SnS = Sn^{2+} + S^{2-}$	1×10^{-26}	26.0
Zinc sulfide	$ZnS = Zn^{2+} + S^{2-}$	1.6×10^{-23}	22.80
Sulfites			
Barium sulfite	$BaSO_3 = Ba^{2+} + SO_3^{2-}$	1.0 x 10-s	8.00
Calcium sulfite	$CaSO_3 = Ca^{2+} + SO_3^{2-}$	1.0×10^{-4}	4.00

สารประกอบ	สมการที่สมดุล	K_{sp}	pK_{sp}
Sulfites (ต่อ)			
Mercury (I) sulfite	$Hg_2SO_3 = Hg_2^{2+} + SO_3^{2-}$	9×10^{-28}	27.0
Silver sulfite	$Ag_2SO_3 = 2Ag^+ + SO,^{2-}$	1.9×10^{-11}	10.72
Strontium sulfite	$SrSO_3 = Sr^{2+} + SO_3^{2}$	3.9×10^{-8}	7.4 1
Thlocyanates			
Copper (I) thiocyanate	$CuSCN = Cu^+ + SCN^-$	4×10^{-14}	13.4
Mercury (I)	$Hg(SCN)_2 = Hg_2^{2+} + 2SCN^-$	3×10^{-20}	19.5
thiocyanate	$AgSCN = Ag^+ + SCN^-$	1×10^{-12}	12.0
Silver thiocyanate			

ตาราง 4B : ค่าคงที่ของการรวมตัวของสารเชิงซ้อน (Formation constants)

ลิแกนด์	สมการที่สมดุล	K_{inst}	pK_{inst}
Ammonia	$[Cd(NH_3)_4]^{2+} = Cd^{2+} + 4NH_3$	1.9×10^{-7}	6.72
	$[Co(NH_3)_6]^{3+} = Co^{3+} + 6NH_3$	2.2×10^{-34}	33.66
	$[Co(NH_3)_4]^{2+} = Co^{2+} + 4NH_3$	9×10^{-6}	5.0
	$[Cu(NH_3)_6]^{2+} = Co^{2+} + 6NH_3$	1.3×10^{-5}	4.89
	$[Cu(NH_3)_4]^{2+} = Cu^{2+} + 4NH_3$	1×10^{-12}	12.0
	$[Ni(NH_3)_4]^{2+} = Ni^{2+} + 4NH_3$	1×10^{-8}	8.0
	$[Ni(NH_3)_6]^{2+} = Ni^{2+} + 6NH_3$	6 x'l 0^{-9}	8.2
	$[Zn(NH_3)_4]^{2+} = Zn^{2+} + 4NH_3$	2×10^{-9}	8.7 (ที่ $30^{\circ}C$)
	$[Ag(NH_3)]^{+} = Ag^{+} + NH,$	5.0×10^{-4}	3.30
	$[Ag(NH_3)_2]^{+} = Ag^{+} + 2NH_3$	5.9×10^{-8}	7.23
Bromide ion	$[HgBr_4]^{2-} = Hg^{2+} + 4Br^{-}$	1×10^{-21}	21.0
Chloride ion	$[CdCl_4]^{2-} = Cd^{2+} + 4Cl^{-}$	9.1×10^{-4}	3.04
	$[FeCl]^{2+} = Fe^{3+} + Cl^{-}$	$3x\ 10^{-2}$	1.5
	$[FeCl_2] = Fe^{3+} + 2Cl^{-}$	0.222	0.654
	$[HgCl_4]^{2-} = Hg^{2+} + 4Cl^{-}$	1×10^{-15}	15.0
	$[AgCl_2]^{-} = Ag^{+} + 2Cl^{-}$	7×10^{-6}	5.2
Cyanide ion	$[Cd(CN)_4]^{2-} = Cd^{2+} + 4CN^{-}$	8×10^{-18}	17.1
	$[Cu(CN)_3]^{2-} = Cu^{+} + 3CN^{-}$	5×10^{-28}	27.3
	$[Cu(CN)_4]^{3-} = Cu^{+} + 4CN^{-}$	5×10^{-30}	29.3
	$[Fe(CN)_6]^{3-} = Fe^{3+} + 6CN^{-}$	1×10^{-42}	42.0
	$[Fe(CN)_6]^{4-} = Fe^{2+} + 6CN^{-}$	1×10^{-35}	35.0
	$[Hg(CN)_4]^{2-} = Hg^{2+} + 4CN^{-}$	3×10^{-42}	41.5
	$[Ni(CN)_4]^{2-} = Ni^{2+} + 4CN^{-}$	1×10^{-22}	22.0
	$[Ag(CN)_2]^{-} = Ag^{+} + 2CN^{-}$	1×10^{-20}	20.0
	$[Zn(CN)_4]^{2-} = Zn^{2+} + 4CN^{-}$	1×10^{-19}	19.0

ลิแกนด์	สมการที่สมดุล	K_{inst}	pK_{inst}
Ethylenediamine	$[CaY]^{2-} = Ca^{2+} + Y^{4-}$	2×10^{-11}	10.7
tetraacetate ion ($=Y^{4-}$)	$[FeY]^{-} = Fe^{3+} + Y^{4-}$	1×10^{-25}	25.0
(ที่ 20° C)	$[MgY]^{2-} = Mg^{2+} + Y^{4-}$	2×10^{-9}	8.7
	$[ZnY]^{2-} = Zn^{2+} + Y^{4-}$	3.1×10^{-17}	16.51
Fluoride ion	$[AlF_6]^{3-} = Al^{3+} + 6F^{-}$	2×10^{-21}	20.7
	$[FeF_6]^{3-} = Fe^{3+} + 6F^{-}$	1×10^{-16}	16.0
Hydroxide ion	$[Al(OH)_4]^{-} = Al^{3+} + 4OH^{-}$	1.2×10^{-34}	33.92
	$[Zn(OH)_4]^{2-} = Zn^{2+} + 4OH^{-}$	2.5×10^{-15}	14.60
Iodide ion	$[CdI_4]^{2-} = Cd^{2+} + 4I^{-}$	7×10^{-7}	6.2
	$[HgI_4]^{2-} = Hg^{2+} + 4I^{-}$	5.3×10^{-31}	30.28
Oxaiate ion	$[Al(C_2O_4)_3]^{3-} = Al^{3+} + 3C_2O_4^{2-}$	5×10^{-17}	16.3
	$[Fe(C_2O_4)_3]^{3-} = Fe^{3+} + 3C_2O_4^{2-}$	6×10^{-21}	20.2
Sulfide ion	$[HgS_2]^{2-} = Hg^{2+} + 2S^{2-}$	2.0×10^{-55}	54.70
sulfite ion	$[Ag(SO_3)_2]^{3-} = Ag^{+} + 2SO_3^{2-}$	3×10^{-9}	8.5
Thiocyanate ion	$[Fe(SCN)]^{2+} = Fe^{3+} + SCN^{-}$	9.4×10^{-4}	3.03
	$[Fe(SCN)_6]^{3-} = Fe^{3+} + 6SCN^{-}$	1×10^{-4}	4.0
	$[Hg(SCN)_4]^{2-} = Hg^{2+} + 4SCN^{-}$	1×10^{-22}	22.0
Thiosulphate ion	$[Ag(S_2O_3)_2]^{3-} = Ag^{+} + 2S_2O_3^{2-}$	1×10^{-13}	13.0

ตาราง 5 B : ค่าคงที่ของการรวมตัวของโลหะ กับ EDTA

(Metal -EDTA Formation constants)

โลหะ	K_{MY}	log K_{MY}
Ag^{+}	2.1 x 10'	7.32
Al^{3+}	1.3 x 10^{16}	16.11
Ba^{2+}	5.8 x 10'	7.76
Bi^{3+}	6.3 x 10^{27}	27.80
Ca^{2+}	5.0 x 10^{10}	10.70
Cd^{2+}	2.9 x 10^{16}	16.46
Co^{2+}	2.0 x 10^{16}	16.30
Cu^{2+}	6.3 x 10^{18}	18.80
Fe^{2+}	2.1 x 10^{14}	14.32
Fe^{3+}	1.3 x 10^{25}	25.11
Hg^{2+}	6.3 x 10^{21}	21.80
Mg^{2+}	4.9 x 10^{8}	8.69 .
Mn^{2+}	6.2 x 10^{13}	13.79
Ni^{2+}	4.2 x 10^{18}	18.62
Pb^{2+}	1.1 x 10^{18}	18.04
Sn^{2+}	2.0 x 10^{18}	18.30
Sr^{2+}	4.3 x 10^{8}	8.63
Zn^{2+}	3.2 x 10^{16}	16.51